

The p-Block Elements Class 11 Notes

Chemistry Chapter 11

• p-Block Elements

Elements belonging to groups 13 to 18 of the periodic table are called p-block elements. General electronic configuration: $ns^2 np^{1-6}$ (except for He)

Table 11.1 General Electronic Configuration and Oxidation States of p-Block Elements

Group	13	14	15	16	17	18
General electronic configuration	$ns^2 np^1$	$ns^2 np^2$	$ns^2 np^3$	$ns^2 np^4$	$ns^2 np^5$	$ns^2 np^6$ ($1s^2$ for He)
First member of the group	B	C	N	O	F	He
Group oxidation state	+3	+4	+5	+6	+7	+8
Other oxidation states	+1	+2, -4	+3, -3	+4, +2, -2	+5, +3, +1, -1	+6, +4, +2

• Group 13 Elements: The Boron Family

Outer Electronic Configuration: $ns^2 np^1$

Atomic Radii: The atomic and ionic radii of group 13 elements are smaller than the corresponding elements of alkali and alkaline earth metals.

Reason: On moving from left to right in a period the effective nuclear charge increases and the outer electrons are pulled more strongly towards the nucleus. This results in decrease in atomic size.

On moving down the group, both atomic and ionic radii expected to increase due to the addition of a new electron shell with each succeeding element.

Exception: Atomic radius of Ga is less than that of Al due to the presence of poor shielding 10d-electrons in gallium.

Ionisation enthalpies: First ionisation enthalpies of the elements of group-13 are less than those of the elements present in group-2 in the same period.

Reason: The removal of p-electron is much easier than the s-electron and therefore, the first ionisation enthalpies ($\Delta_i H_1$) of the elements of group 13 are lower as compared to the corresponding elements of group 2.

On moving down the group 13 from B to Al the first-ionization enthalpies ($\Delta_i H_1$) decrease due to an increase in atomic size and screening effect which outweigh the effect of increased nuclear charge.

There is discontinuity expected in the ionisation enthalpy values between Al and Ga and between In and Tl due to enability of d- and f-electrons which have low screening effect to compensate the increase in nuclear charge.

Electronegativity: Down the group, electronegativity first decreases from B to Al and then increases.

This is due to discrepancies in the atomic size of the elements.

Physical Properties

(i) Due to strong crystalline lattice boron has high melting point. Rest of the members of this family have low melting point.

- (ii) Boron is extremely hard and black coloured solid and non metallic in nature.
 (iii) Other members of this family are soft metals with low melting point and high electrical conductivity.

Chemical Properties

Oxidation states: The first two elements boron and aluminium show only +3 oxidation state ~ in the compounds but the other elements of this group gallium, indium and thalium also exhibit +1 oxidation state in addition to +3 oxidation state i.e., they show variable oxidation states.

As we move down the group, the stability of +3 oxidation state decreases while that of +1 oxidation state progressively increases.

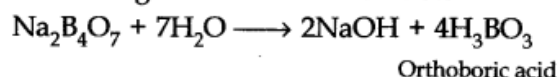
• Some Compounds of Boron

(i) **Borax:** Borax is a white crystalline solid.

It contains tetranuclear units. $[B_4O_5(OH)_4]^{2-}$

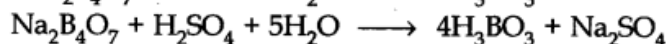
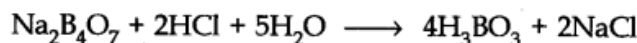
Its formula is $Na_2[B_4O_5(OH)_4] \cdot 8H_2O$

Borax dissolves in water to give an alkaline solution

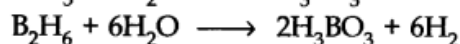
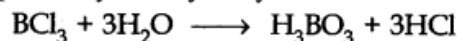


Borax is used as a water softner and cleaning agent.

(ii) **Orthoboric acid (H_3BO_3) or $B(OH)_3$:** Boric acid can be prepared by the acidification of aqueous solution of borax.



It can also be prepared by the hydrolysis of boron compounds.



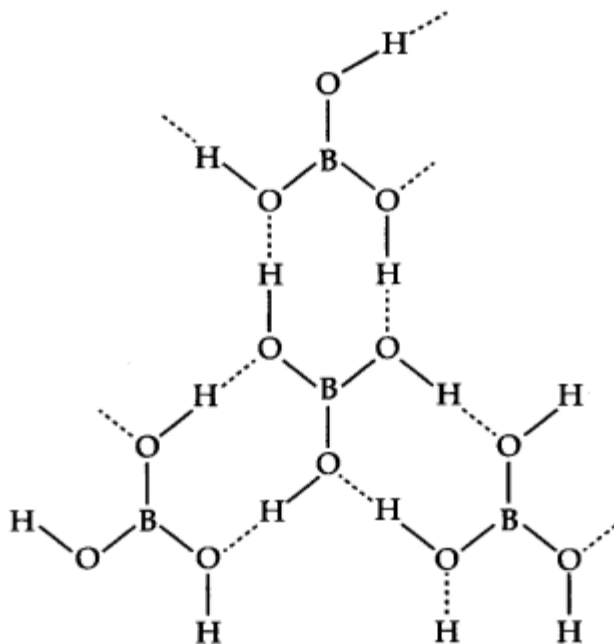
Physical properties of boric acid:

- (i) It is a white crystalline solid.
 (ii) It is soft soapy in touch.
 (iii) It is sparingly soluble in cold water but fairly soluble in hot water.

Uses:

- (i) In the manufacture of heat resistant borosilicate glazes.
 (ii) As a preservative for milk and food stuffs.
 (iii) In the manufacture of enamels and glazes in pottery.

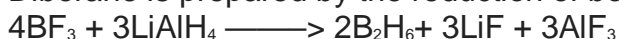
Structure of boric acid



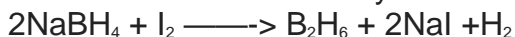
Structure of boric acid; the dotted lines represent hydrogen bonds

(iii) Diborane, (B_2H_6): The series of compounds of boron with hydrogen is known as boranes.

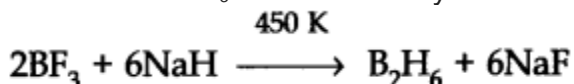
Diborane is prepared by the reduction of boron trifluoride with $LiAlH_4$ in diethyl ether.



Laboratory method of preparation. In laboratory diborane is prepared by the oxidation of sodium borohydride with iodine.



Industrial method of preparation. On industrial scale, diborane is prepared by reduction of BF_3 with sodium hydride.

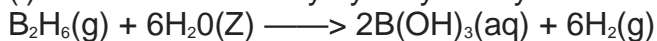


Physical Properties:

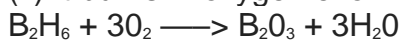
- (i) Diborane is a colourless, highly toxic gas with a b.p. of 180 K.
- (ii) Diborane catches fire spontaneously upon exposure to air.
- (iii) Higher boranes are spontaneously flammable in air.

Chemical properties:

(i) Boranes are readily hydrolysed by water to form boric acid

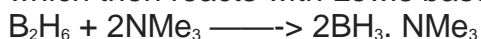


(ii) It burns in oxygen evolving an enormous amount of heat



(iii) Reaction with Lewis base:

Diborane on treatment with lewis bases undergo cleavage reactions to form borane which then reacts with Lewis bases to form adducts.



• **Group 14 Elements:** The Carbon Family Group 14 includes carbon (C), silicon (Si), Germanium (Ge), tin (Sn) and lead (Pb).

General electronic configuration of carbon family is ns^2np^2 .

Carbon: Carbon is the seventeenth most abundant element by weight in the earth's

crust.

(i) It is available as coal, graphite and diamond. In combined state it is present in metal carbonates, hydrocarbons and carbon dioxide gas (0.03%) in air.

(ii) Naturally occurring carbon contains two stable isotopes ^{12}C and ^{13}C and third isotope ^{14}C . ^{14}C is a radioactive isotope with half life 5770 years and is used for radiocarbon dating.

Covalent radius: Covalent radius expected to increase from C to Si. From Si to Pb small increase is found.

Reason: Due to the addition of a new energy shell in each succeeding element. The increase in covalent radii from Si to Pb is small due to ineffective shielding of the valence electrons by the intervening d- and f orbitals.

Ionization Enthalpy: The first ionization enthalpies of group 14 elements are higher than those of the corresponding group 13 elements.

Reason: Because effective nuclear charge increases and size of the atoms becomes smaller. First ionization enthalpy decreases on moving down the group from carbon to tin.

The decrease is very sharp from carbon to silicon while there is slight increase in the first ionization enthalpy of lead as compared to that of tin.

Electronegativity: Group 14 elements are smaller in size as compared to group 13 elements that's why this group are slightly more electronegative than group 13. From Si to Pb it is almost same. Small increase in ionization enthalpy from Sn to Pb is due to the effect of increased nuclear charge outweighs the shielding effect due to the presence of additional 4f- and 5d-electrons.

Physical properties:

(i) All the elements of group 14 elements are solids. They are less metallic than group 13.

(ii) M.P. and boiling points of group 14 elements are generally high.

Chemical properties:

Carbon and silicon mostly show +4 oxidation state. Germanium forms stable compounds in +4 state and only few compounds in +2 state.

Tin forms compounds in both oxidation states. Lead forms compounds in +2 state are stable and in +4 state are strong oxidising agents.

• **Anomalous Behaviour of Carbon**

Carbon, differs from the rest of the member of its family. The main reason for the anomalous behaviour is:

(i) exceptionally small atomic and ionic size

(ii) higher ionization enthalpy

(iii) absence of d-orbitals in the valence shell.

(iv) Higher electronegativity.

It can be explained as follows:

=> Since carbon has only s and p-orbitals it can accommodate only four pairs of electrons ; other member can expand their covalence due to the presence of d-orbitals.

=> Carbon can form $\text{P}\pi\text{-P}\pi$ multiple bonds with itself and other atoms having small size and high electronegativity.

For example, $\text{C}=\text{C}$, $\text{C}\equiv\text{C}$, $\text{C}=\text{O}$, $\text{C}=\text{S}$ and $\text{C}\equiv\text{N}$

Heavier elements do not form $\text{P}\pi\text{-P}\pi$ bonds because their atomic orbitals are too large and diffuse to have effective overlapping.

=> Carbon atoms have the tendency to link with one another through covalent bonds to form chains and rings. This property is called catenation.

Down the group property to show catenation decreases.

The order of catenation is $C \gg Si > Ge \approx Sn$

Lead does not show catenation.

• Allotropes of Carbon

The property of an element to exist in two or more forms which have different physical properties but identical chemical properties is called allotropy and different forms are called allotropes. Carbon exists in two allotropic forms:

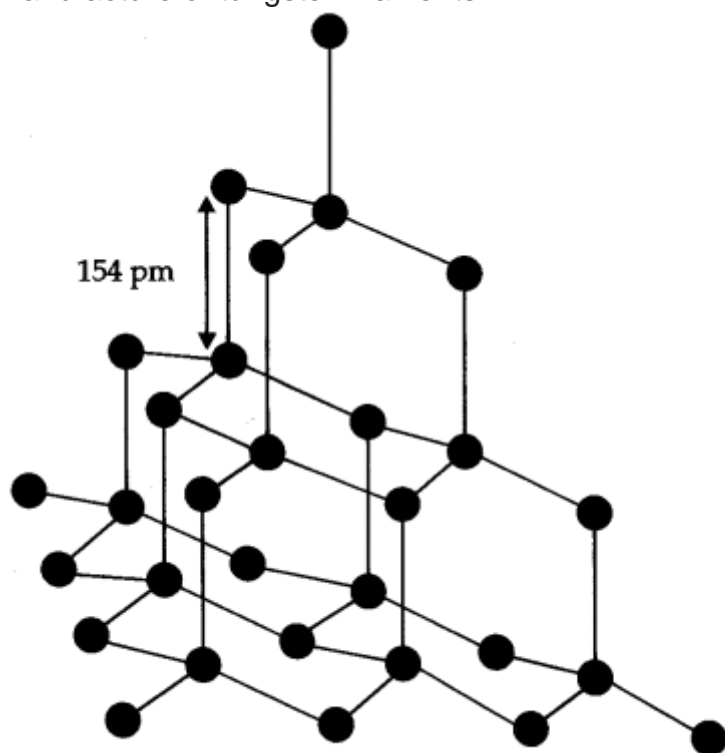
(i) Crystalline (ii) Amorphous

Crystalline form of carbon: Diamond, Graphite, Fullerenes
Diamond: In diamond each carbon atom undergoes sp^3 hybridisation. Each carbon is tetrahedrally linked to four other carbon atoms. The C—C bond length is 154 pm.

Properties:

(i) It is the hardest substance on earth.

(ii) It is used as an abrasive for sharpening hard tools in making dyes and in manufacture of tungsten filaments.



The structure of diamond

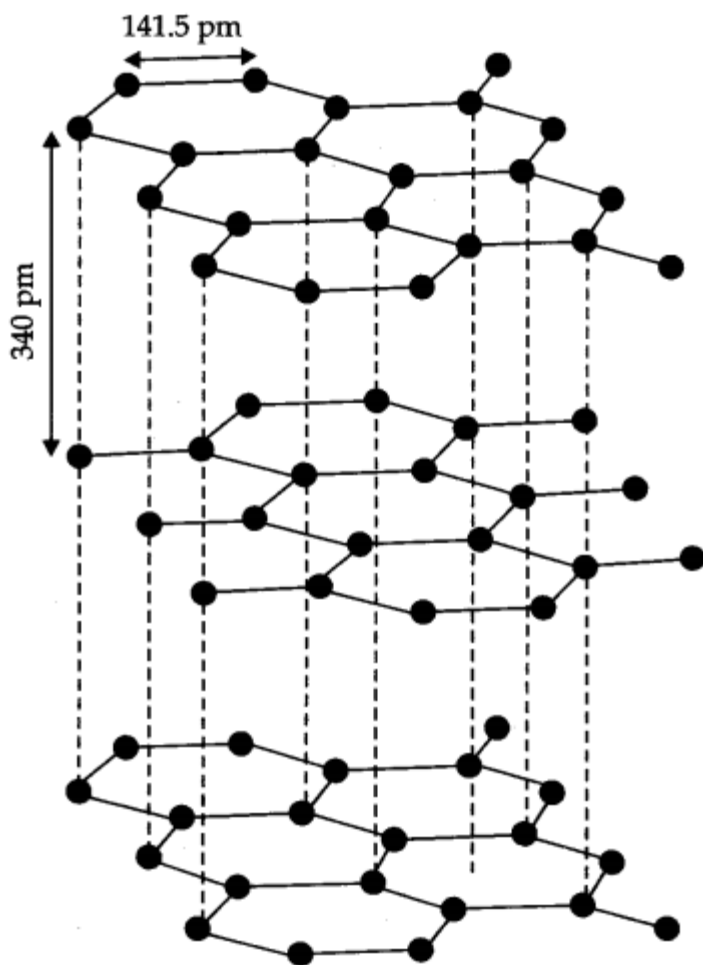
Graphite: In graphite, carbon is sp^2 -hybridized. Graphite has a two-dimensional sheet like structure consisting of a number of hexagonal rings fused together. Layers are held by van der Waals forces and distance between two layers is 340 pm.

Properties:

(i) Graphite conducts electricity along the sheet.

(ii) It is very soft and slippery.

(iii) Used as a dry lubricant in machines running at high temperature, where oil cannot be used as a lubricant.



The structure of graphite

Fullerenes: Fullerenes was discovered collectively by three scientists namely E. Smalley, R.F. Curl and H.W. Kroto.

Preparation:

Fullerenes is prepared by heating of graphite in an electric arc in the presence of inert gas such as helium or argon.

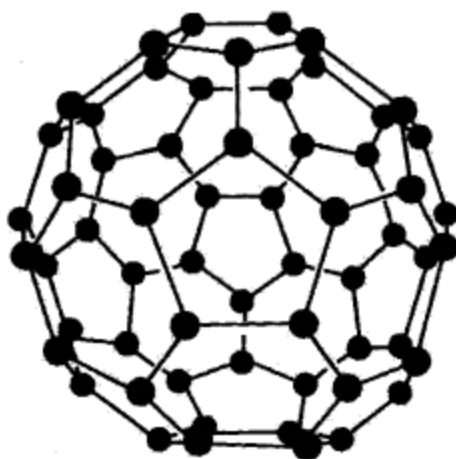
The sooty material formed by the condensation of vapourised C^n small molecules consists of mainly with smaller quantity of C_{70} and traces of other fullerenes consisting of even number of carbon atoms up to 350 or above.

Fullerenes are cage like molecules. C_{60} molecule has a shape like soccer ball and called Buckminsterfullerenes. It is the most stable.

It contains 20 six-membered rings and 12 five-membered rings.

Six-membered rings are fused to both the other six-membered rings and five-membered rings but the five-membered rings are connected only to six-membered rings.

All the carbon atoms are equal and they undergo sp^2 -Hybridization.



The structure of C_{60} Buckminsterfullerene:

Note that molecule has the shape of a soccer ball (football).

Properties:

- (i) Fullerenes being covalent are soluble in organic solvents.
- (ii) It also forms platinum complexes.

Amorphous allotropic forms of carbon coke: It is a greyish black hard solid and is obtained by destructive distillation.

Wood charcoal: It is obtained by strong heating of wood in a limited supply of air.

Animal charcoal: It is obtained by the destructive distillation of bones.

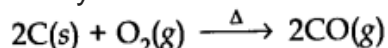
Uses of carbon:

- (i) Graphite fibre are used for making superior sports goods such as tennis and badminton rackets, fishing rods.
- (ii) Being good conductor graphite is used for making electrodes for batteries and industrial electrolysis.
- (iii) Being highly porous, activated charcoal is used for absorbing poisonous gases in gas masks. It is used to decolourize sugar.
- (iv) Carbon black is used as black pigment in black ink and as filler in automobile tyres.
- (v) Coke is extensively used as reducing agent in metallurgy.
- (vi) Diamond is a precious stone.

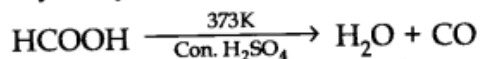
• Some Important Compounds of Carbon and Silicon

Carbon Monoxide

Preparation: It is prepared by direct oxidation of C in limited supply of oxygen.



On small scale it is prepared by dehydration of formic acid with Con. H_2SO_4 at 373 K.



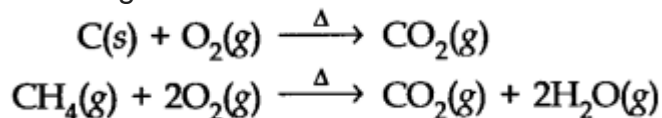
Properties:

- (i) Carbon monoxide is a colourless and odourless gas.
 - (ii) It is almost insoluble in water.
 - (iii) It is powerful reducing agent and reduces almost all metal oxides except alkali and alkaline earth metal oxides.
 - (iv) In CO molecule there are one σ (sigma) and two π bonds between carbon and oxygen.
- : C = O :
- (v) It is highly porous in nature. It forms a complex with haemoglobin which is about

300 times more stable than the oxygen-haemoglobin complex. This prevents haemoglobin in the red blood corpuscles from carrying oxygen round the body, thereby causing suffocation ultimately leading to death.

Carbon Dioxide

Preparation: It is prepared by complete combustion of carbon and carbon containing fuels in



Properties:

- (i) It is a colourless and odourless gas.
- (ii) It is slightly soluble in water. When CO_2 dissolves in water only some of the molecules react with water to form carbonic acid.
- (iii) It is not poisonous like CO.

But increase in combustion of fossil fuels and decomposition of limestone for cement manufacture increase of CO_2 in the atmosphere is one of the main reasons of green house effect.

Silicon dioxide (SiO_2)

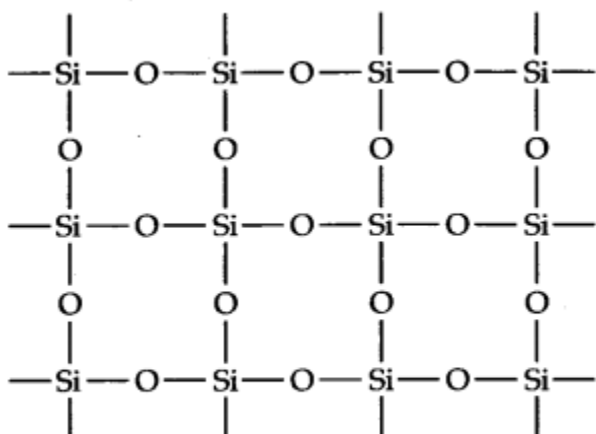
Silicon dioxide, commonly known as silica, occurs in various crystallographic forms. For example, Quartz, Cristobalite and stishovite are some of the crystalline forms of silica.

Structure:

Silicon dioxide is a covalent three dimensional network solid.

Each silicon atom is covalently bonded in a tetrahedral manner to four oxygen atoms.

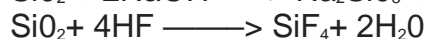
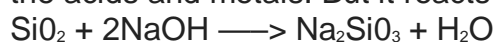
Each oxygen atom in turn covalently bonded to another silicon atoms as shown below:



Three dimensional structure of SiO_2

Properties:

- (i) In normal form silica is very less reactive.
- (ii) At elevated temperature it does not react with halogens, dihydrogen and most of the acids and metals. But it reacts with HF and NaOH.



Uses:

- (i) Quartz is extensively used as a piezoelectric material.

- (ii) Silica gel is used as adsorbent in chromatography.
(iii) An amorphous form of silica, kieselghur is used in filtration plants.

• **P-Block elements:** Contains, metals, non-metals and metalloids.

• **General configuration:** ns^2np^{1-6}

– Boron is a typical non-metal and the other members are metals.

– Boron halides are considered to behave like Lewis acids.

– Boric acid is a Lewis acid.

– Borax is a white crystalline solid formula is $Na_2 [B_4O_5(OH)_4] \cdot 8H_2O$

– Aluminium exhibits +3 oxidation state.

– **Catenation:** Ability of carbon to form chains or rings not only with C-C single bonds but also with multiple bonds ($C = C$ or $C \equiv C$). The tendency of catenation decreases as $C \gg Si > Ge \approx Sn$

– Allotropy: The important allotropes of carbon are diamond, graphite, and fullerenes.

– The members of carbon family exhibit +4 and +2 oxidation state. The tendency to show +2 oxidation state increases among heavier elements.

– Lead in +2 state is stable whereas in +4 oxidation state it is a strong oxidising agent.

– Carbon monoxide is neutral whereas CO_2 is acidic in nature.

– Carbon monoxide having lone pair of electrons on C forms metal carbonyls.

– Carbon monoxide forms a haemoglobin complex which is deadly poisonous due to its higher stability.

– Zeolites are complex aluminium silicates.